Lauren Dillingham

Copper Plating

A Homeschooler's Experiment from Theo Gray's Mad Science

Coating metal objects with copper using electricity and a chemical bath…

I took a piece of copper and a chemical bath, **E.B ULTRA CU RTP**, provided by **EPI Electro-Chemical Products, Inc.**, and ripped electrons from a copper anode through the processes of electroplating. By attaching the positive end of a power source to the anode, and the negative end to the metal ladle, the ions traveled through the bath, regained electrons at the surface of the ladle, where they turned back into solid metal. The result was (Ta-Da!) a shiny copper coating in just two or three minutes.

Experiments you can do at home, but probably shouldn't!

Materials Needed:

- Fine Sandpaper
- Power supply
- E.B. ULTRA CU RTP
- Copper anode
- Alligator Clips
- Metal Object
- A glass bowl/Tank

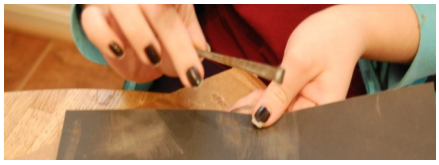
March 1, 2010

Homers Shine!

29 Cu
Copper

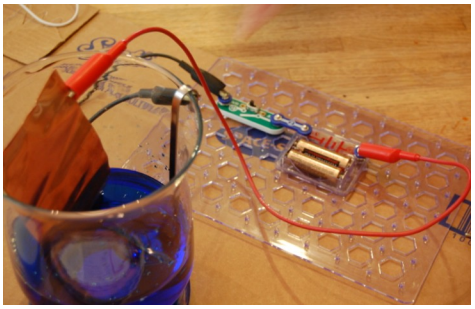
Melting Point: 1,984 F
Boiling Point: 4,643 F

Found in: bloodstream of nearly all plants and animals.
Uses: Electrical wiring, plumbing,

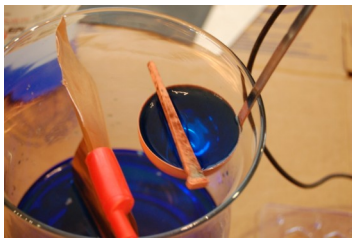


Preparing the metal object.

Placing the object into the bath.



Zapping the anode—Zap Time!!!



Ta-Da!
Transformation Complete!